U.S. CHEMICAL SAFETY AND

HAZARD INVESTIGATION BOARD

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T2 Laboratories Explosion and Fire

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Public Meeting Jacksonville, Florida

Marriott Hotel Jacksonville Florida Room September 15, 2009

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6:00 p.m.

PRESENT:

JOHN BRESLAND, Chairman

GARY VISSCHER, Board Member

WILLIAM WARK, Board Member

WILLIAM WRIGHT, Board Member

CHRIS WARNER, General Counsel

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2	(6:00 p.m.)	
3	CHAIRMAN BRESLAND: We'll be	
4	starting in about 30 seconds so if I could ask	
5	everybody to take a seat and we'll get going.	
б	That's the official gavel. Good	
7	evening and welcome to the public meeting of	
8	the U.S. Chemical Safety Board or as we know	
9	it the CSB. I'm John Bresland, Chairman and	
10	CEO of the board. With me today are board	
11	members Gary Visscher on my right, William	
12	Wark and William Wright on my left. Also	
13	joining us is our General Counsel, Chris	
14	Warner, and CSB staff members whose efforts	
15	have facilitated this evening's meeting.	
16	The CSB is an independent	
17	non-regulatory federal agency that	
18	investigates major chemical accidents of fixed	
19	facilities.	
20	Our investigation examines all	
21	aspects of chemical accidents including	
22	physical causes related to equipment design as	

well as inadequacies in regulations, industry 1 2 standards and safety management systems. Ultimately we issue safety 3 recommendations which are designed to prevent 4 5 similar accidents in the future. 6 The purpose for this evening's 7 meeting is to allow the CSB investigative team to present their findings. The board members 8 9 will then vote on the final report into the December 19th, 2007 explosion and fire at the 10 T2 Laboratories in Jacksonville, Florida. 11 12 Following that presentation we 13 will open the floor to comments from members of the public who are here this evening. The 14 meeting will conclude with a discussion by the 15 board and the vote on the final report. 16 Before we begin I'd like to point 17 out some safety information. Please take a 18 moment to note the locations of the exits from 19 this meeting room. The exits are at the back 20 of the room and over here and lead to the 21 22 outside in case of an emergency.

1	I also ask that you please mute
2	your cell phones so that these proceedings are
3	not disturbed. I'll pause and give you a
4	second to mute your cell phones. Thank you.
5	At 1:30 p.m. on December 19th,
6	2007 an explosion destroyed the T2
7	Laboratories facility in Jacksonville. A
8	runaway reaction occurred during the
9	production of methylcyclopentadienyl manganese
10	tricarbonyl, or MCMT as we'll refer to it this
11	evening. It resulted in an explosion and
12	fire. It killed four employees and injured
13	four others. It also injured 28 members of
14	the public.
15	The incident at T2 Laboratories
16	included one of the most powerful explosions
17	that the CSB has ever investigated, a blast
18	that was felt 15 miles away in downtown
19	Jacksonville. Debris from the explosion was
20	found up to one mile away. The blast damaged
21	buildings within one quarter mile of the
22	facility. This was a tragic unnecessary loss

of life and the CSB immediately began an
 investigation.

Many of you here this evening were 3 affected by the devastation that resulted from 4 5 this incident. To those of you in the audience who lost friends, family members or were 6 7 yourself injured by the incident, please allow me to extend my deepest sympathy. 8 9 Let me explain the process the board follows to complete an investigation: 10 Each independent board member, the 11 four board members who are here this evening 12 13 have had an opportunity to study the draft report. At this public meeting we will hear 14 a presentation by the investigation team. 15 The board members may direct questions about the 16 investigation to the team. There will then be 17 a board member vote on the report and its 18 recommendations. 19 20 Our objective is to leave here this evening with strong effective 21 recommendations based on the report's 22

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1 findings. The board's goal is that the report 2 and it's recommendations will help prevent, 3 and I emphasize prevent, similar accidents in 4 the future.

5 If anyone in the audience wishes to comment publicly after the investigations 6 7 -- the investigators' presentation, please sign up at the table in the check-in area and 8 9 I will call your name at the appropriate time. 10 However, if you haven't signed up we certainly will offer the opportunity for additional 11 people to make comments also. Please note 12 13 that we will have to limit the public comments to three minutes each. 14

Also note that we are not able to 15 take questions for the investigators directly 16 from the audience so I'll ask that all 17 questions or comments be directed to me as the 18 presiding official. If there's a point that 19 20 it raised in your comment where I believe the investigation staff can provide some immediate 21 clarification, I will ask them to do so. 22

1	I would like to thank the team for
2	their diligent work on this investigation.
3	I will now recognize any other
4	board members for an opening statement.
5	Mr. Visscher?
6	MR. VISSCHER: I have none.
7	CHAIRMAN BRESLAND: Mr. Wark?
8	MR. WARK: I have none.
9	CHAIRMAN BRESLAND: Mr. Wright?
10	MR. WRIGHT: I have none.
11	CHAIRMAN BRESLAND: Thank you.
12	At this time I'll ask the
13	investigation supervisor Robert Hall to
14	introduce the investigation team.
15	Mr. Hall is a Registered
16	Professional Engineer with a bachelor's degree
17	in nuclear engineering and a master's degree
18	in mechanical engineering. He has more than
19	25 years experience in the design, inspection
20	and evaluation of hazardous systems for the
21	chemical, nuclear and the aerospace
22	industries. Mr. Hall served as the

investigator in charge for this investigation. 1 2 Now let me turn it over to Mr. Hall. MR. HALL: Thank you, Mr. 3 Bresland. 4 5 With me tonight is the team that helped me investigate this accident. 6 On the 7 far end of the table Mary Nikityn. She holds a bachelor's degree in psychology from 8 9 Southern Adventist University and is currently a doctoral student in clinical psychology at 10 George Washington University. Prior to 11 joining the CSB Ms. Nikityn worked as a 12 13 regulatory specialist in the compliance training industry with expertise in EPA 14 hazardous waste and DOT hazardous materials 15 regulation. 16 Next to Mary is Lucy Sciallo. 17 She is a graduate out of Pennsylvania State 18 University with a Bachelor of Science degree 19 20 in Industrial Health and Safety. Prior to joining CSB Lucy worked in the oil industry as 21

22 a health and safety specialist. She holds a

Graduate Safety Practioner designation from 1 2 the Board of Certified Safety Professionals. And lastly Mr. Jeffrey Wanko. 3 Mr. Wanko received a bachelor's degree in chemical 4 5 engineering from Syracuse University and a master's from the Illinois Institute of 6 7 Technology. Mr. Wanko has been a safety professional for more than 16 years 8 9 specializing in process safety in the pharmaceutical, specialty chemical and food 10 Mr. Wanko is a licensed Professional 11 sectors. Engineer and a Certified Safety Professional. 12 13 Chairman Bresland, members of the board, Chief Counsel, ladies and gentlemen, 14 good evening. 15 The investigation team is pleased 16 to present the facts, findings and causes of 17 the runaway reaction, explosion and fire that 18 occurred at T2 Laboratories in Jacksonville on 19 December 19th, 2007. 20 The incident at T2 laboratories 21 22 included one of the most powerful explosions

1 that the CSB has ever investigated, a blast that was felt 15 miles away in downtown 2 Jacksonville. The explosion caused wide 3 spread community damage, flinging debris in 4 5 all directions. Four T2 employees were T2 employees and many members of the 6 skilled. 7 public working at nearby businesses were injured. Dozens of firefighters spent many 8 9 hours battling the chemical fire. Businesses 10 were damaged. Buildings had to be demolished and rebuilt. And the environmental clean up 11 is still ongoing. 12

13 The impact of this incident on the 14 community was severe and the CSB has spent many months researching and analyzing the 15 incident in the hope of preventing future such 16 Tonight we are presenting our 17 occurrences. findings and recommendations to both the board 18 and the Jacksonville community. 19 20 The presentation includes a 21 summary of the incident, a computer 22 reenactment of the most likely scenario as

well as the presentation of the team's 1 2 findings and recommendations. Following the presentation we'll have an opportunity for the 3 public to comment and the board will vote on 4 5 the team's report and recommendations. Before I discuss the incident I'd 6 7 like to give you a very brief background on 2T Laboratories. 8 9 T2 Laboratories, which is no 10 longer in business, was a chemical manufacturing facility located on the north 11 side of Jacksonville. T2 began as a specialty 12 13 solvent blending company in 1996. In 2000 T2 began research and development on a 14 manufacturing process for a chemical called 15 methylcyclopentadienyl manganese tricarbonyl 16 17 or MCMT. 18 Over the next few years T2 designed and built the chemical process line 19 20 and began producing MCMT at the Faye Road site. By 2007 at the time of the incident T2 21 22 had 12 employees and had manufactured 174

1 batches of MCMT. T2 produced MCMT from 2004 2 until the incident in 2007 selling it under the tradename Ecotane. 3 MCMT is used as an octane booster 4 5 in gasoline and is sold in the U.S. primarily as a specialty product in automotive stores. 6 7 T2 sold MCMT to a U.S. distribution company and other companies 8 9 internationally. MCMT is highly toxic through 10 inhalation and skin absorption. It is regulated by the Environmental Protection 11 12 Agency as an extremely hazardous substance. 13 T2 manufactured MCMT in a single reactor in batches using a three-step chemical process. 14 On the morning of December 19th, 15 2007 T2 began producing its 175th batch. 16 The reaction proceeded normally throughout the 17 morning; however, shortly after 1:00 p.m. a 18 cooling problem was reported. The reaction 19 20 temperature began to rise past the normal 21 operation point and the reaction proceeded 22 uncontrollably. As temperature and pressure

rapidly increased, the reactor burst and its 1 2 contents exploded in a blast equivalent to 1400 pounds of TNT. Video footage shows the 3 4 very energetic nature of this explosion. 5 This first clip was taken by an infrared surveillance camera located several 6 7 miles away. The two large cooling towers that you see in the center of the screen are 460 8 9 feet high. As the heat from the blast 10 dissipates, you can see smaller objects These are like pieces of sodium 11 twinkling. metal from inside the reactor reacting with 12 13 the moisture in the air as they fall. You can also see a smokestack in 14 line with the explosion. This smokestack is 15 640 feet high. 16 This business is located about 620 17 feet from the reactor structure. The building 18 was damaged and eight of the 16 employees 19 working there were injured. 20 This business is located about 700 21 22 feet away. As you see in the security video,

debris rained down on the workers outside.
 Inside workers looked towards the reactor
 shortly before being showered by glass as the
 windows shattered.

5 This is an aerial photograph of 6 the T2 site the day after the incident. Here 7 in the center you can see the area where the 8 reactor structure stood. Large pieces of the 9 reactor shell were thrown here and here.

10 Another large piece of the reactor 11 weighing about 2000 pounds was found about 400 12 a feet away in this direction. Two steel 13 columns that supported the reactor were found 14 a thousand feet down Faye Road in both 15 directions.

16 This area of the facility was the 17 tank farm and contained thousands of gallons 18 of flammable solvents. After the explosion 19 those ignited and continued to burn for many 20 hours.

21 The control room where two T222 employees were killed was a square 12-foot

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concrete building located here just 50 feet 1 2 from the reactor. As you see the damage on site was devastating. The damage off site and 3 4 the impact to the community were also 5 extensive. The explosion killed four T2 employees and injured 31 of 118 people that 6 7 were working within a thousand feet of the reactor. 8 9 One T2 employee sustained life-threatening injuries that required 10 extensive hospitalization. 11 12 Many of the injuries to people 13 working in nearby businesses were caused by secondary missiles as pieces of structure, 14 office furniture and glass shards were thrown 15 by the blast. 16 Five buildings belonging to 17 businesses on Faye Road sustained so much 18 damage that they were condemned as unsafe 19 structures and had to be rebuilt. 20 Debris flew in all directions throughout a one mile 21 radius. 22

1	This building is near T2. Is it
2	one of those that was condemned as unsafe.
3	The business had to relocate operations while
4	the building was reconstructed.
5	You see here a trailer used by a
6	trucking company adjacent to T2. This
7	structure was destroyed by the blast shortly
8	after trucking company employees had been
9	dismissed for the day. Had anyone been working
10	here, it is likely they would have been
11	severely injured or killed.
12	Many organizations responded to
13	the large explosion, huge fire and extensive
14	damage. The City of Jacksonville Fire and
15	Rescue Department and the sheriff's office
16	managed the incident. While the Environmental
17	Resource Management Division assessed the
18	environmental impact and the Planning and
19	Development Department surveyed damage and
20	condemned unsafe structures.
21	The Naval Air Station Mayport and
22	the Jacksonville Internation Airport Fire

Department helped fight the blaze. 1 The Florida Department of Environmental Protection 2 and the American Red Cross responded to 3 4 oversee site clean up and assist emergency 5 workers respectively. 6 The emergency response to this 7 incident was excellent as many of the agencies cooperated to safely control the fires and 8 9 event further harm. Over 90 firefighters 10 entered the fight to battle the chemical fire. 11 None were injured. 12 Due to the impact of this 13 incident, the CSB decided to deploy a team on December 19th. The team arrived in 14 Jacksonville on the next morning and began the 15 investigation activity but was unable to enter 16 the site for another five weeks because of the 17 highly toxic MCMT and the difficulty of 18 testing for its presence on site. 19 20 While waiting to access the site, the team conducted many investigation 21 The team conducted 48 interviews 22 activities.

with T2 employees and eyewitnesses from other
 businesses. We also examined the reactor
 vessel fragments found on site and collected
 more parts of the reactor from the surrounding
 woods.

6 The team examined and catalogued 7 all of the damaged structures and used this 8 information to determine the strength of the 9 explosion. The CSB submitted dozens of 10 document requests to T2 and other companies 11 and agencies and reviewed thousands of 12 documents received in response.

13 Finally the CSB recreated T2's chemical reaction in a laboratory and 14 duplicated the runaway reaction that occurred 15 on December 19th. A reenactment of what most 16 likely occurred is presented as an animation 17 which is part of a CSB full safety video on T2 18 which will be released soon. I will now 19 20 present the animated reenactment to you. "T2 produced MCMT, a 21 VOICEOVER: 22 gasoline additive in batches using a 2500

gallon reactor. An operator controlled the process with a computerized system in a nearby control room. In the first step liquid chemicals and sodium metal were loaded into the reactor, heated and then mixed with an agitator.

7 "The reaction produced hydrogen which was vented to the atmosphere. In normal 8 9 operations when the temperature reached 300 10 degrees Farenheit the operator would turn off the heating system. But because this reaction 11 was exothermic or heat producing, the 12 13 temperature inside the reactor would continue to rise. 14 "At 360 degrees operators would 15 begin to periodically fill the reactor's 16 cooling jacket with water. As the water 17

18 boiled heat was removed controlling the

19 temperature.

20 "However, on the day of the
21 accident the CSB found that the operator tried
22 to cool the reactor as usual but the cooling

system likely malfunctioned perhaps due to a
 blockage in the water supply piping or a valve
 failure. The temperature and pressure inside
 the reactor began to rise in a runaway
 chemical reaction.

6 "T2's co-owners returned to the 7 plant after a worker called to report the 8 cooling problem. While one owner searched for 9 the plant mechanic the other went to the 10 control room. Concerned about a possible 11 fire, he warned employees to move away from 12 the reactor.

13 "Inside the reactor the pressure was still increasing reaching 400 pounds per 14 square inch and bursting the rupture disk. 15 Witnesses heard a sound like a jet engine as 16 high pressured gas began to vent from the 17 reactor. But it was too late. Within ten 18 seconds there as a massive explosion 19 20 equivalent to about 1400 pounds of TNT." 21 MR. HALL: As you've just seen in 22 the animation, the explosion that occurred at

T2 was very powerful and destroyed the
 facility. The following photos show the same
 views of T2 before and after the incident.
 I will now present the team

5 findings that the team determined from the
6 investigation.

7 Based on our evidence collection, the team identified several key investigation 8 9 areas. First the chemistry used by T2 was 10 exothermic or heat producing. Exothermic chemistry presents specific hazards and 11 requirements for reactor cooling and relief 12 13 system design. Hazard recognition is of particular importance when designing 14 exothermic chemical processes. 15 The team investigated T2's 16 recognition of the reactive hazards as well as 17 agitation on reactive hazards that are 18 commonly provided to undergraduate chemical 19 20 engineers. These ares were key in understanding incident causes and the team 21 22 will present each one in more detail. First

Investigator Wanko will present the results of
 our chemical testing.

Thank you, Mr. Hall. 3 MR. WANKO: 4 T2's owners based the process on a 5 three-step chemical recipe they found in patents dating back to 1950. The patents 6 7 described the different chemicals that could be used to reach the final product, the 8 9 intermediate steps and the processes 10 condition. The patents though do not describe 11 the safety measures that must be in place for full scale production. 12 13 The patents also do not address the heat production or exotherm that occurs in 14 the first step of the process. The first step 15 was about two and a half hours into the 16 production of the batch when the incident 17 occurred. 18 The first stage of the process 19 required initial heating to start the 20

21 reaction. At a predetermined set point

22 heating is shut off as the reaction produces

enough heat to sustain itself. At this point 1 2 the reaction requires periodic cooling to remove heat from the system. 3 4 The team hypothesized that an 5 uncontrolled chemical reaction caused the T2 6 explosion so we recreated the standard 7 chemical recipe of the first step that T2 used and were using on the day of the incident. 8 9 The recipe included 10 methylcyclopentadiene and diethylene glycol dimethyl either, also known as diglyme, and 11 sodium metal. Using only 50 grams of material 12 13 in the same proportions as T2, the team found that the first step exhibits an initial 14 exotherm or heat generating period that can be 15 controlled through cooling and venting hot 16 17 gasses. 18 However, above 390 degrees Fahrenheit, only 30 degrees higher than the 19 20 normal operating temperature, a second undesired reaction occurs that exhibits a 21 22 second and more energetic exotherm.

		Pag
1	If cooling or venting were to have	
2	failed during the first step of the process,	
3	temperature and pressure and therefore	
4	reaction rate would continue to rise.	
5	If conditions are not corrected	
б	quickly a second undesired reaction occurs and	
7	the temperature and pressure rises very	
8	quickly and uncontrollably.	
9	T2's development work was	
10	incapable of determining exothermic activity	
11	and this uncontrollable outcome went	
12	undetected.	
13	This graph illustrates the results	
14	of one of the teams' laboratory run using the	
15	T2 recipe of MCPD, diglyme and metal.	
16	Here in the first desired reaction	
17	temperature and pressure rise gradually over	
18	two and a half hours. At about 390-degrees	
19	Farenheit the second undesirable reaction	
20	begins. Then at about 500 degrees Farenheit	
21	the temperature and pressure rise very	
22	quickly.	

1	In this region the temperature is
2	rising at about 70 degrees Farenheit per
3	second, and pressure at over 500 pounds per
4	square inch per second.
5	Since this chemistry is so
6	energetic we tested a very small amount in
7	this apparatus. It's about half the size of
8	a soda can.
9	Here you can see the test
10	apparatus before and after the second
11	undesired reaction which occurred with less
12	than two ounces of material. The T2's reactor
13	in contrast held over six tons during the
14	first step.
15	With pressure rising over 500
16	pounds per square inch per second, pressure
17	inside the vessel reached 5000 pounds per
18	square inch and burst in less than 10 seconds
19	after the emergency vents opened.
20	Here you can see two large pieces
21	of the reactor. This large piece is from the
22	bottom of the reactor. And this piece is a

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portion of the top weighing about 2000 pounds. 1 It was thrown through a pipe rack, impacted 2 and bent a railroad track and damaged a 3 building about 400 feet away. 4 5 Mr. Hall. 6 MR. HALL: Thank you Investigator 7 Wanko. As you saw in the video, before 8 9 the incident the reactor was held vertically 10 in a tall steel support structure. The reactor itself was about ten feet high and 11 seven feet wide. 12 13 As Investigator Wanko discussed, exothermic reactions generate high 14 temperatures and pressure. To control these 15 16 cooling and pressure release systems are extremely important. Cooling for the reactor 17 was provided by an exterior cooling jacket. 18 City water was piped in from the bottom of the 19 jacket and allowed to boil and vent. 20 The heat required to boil the water was taken out of 21 the reaction. 22

Water was piped into the jacket
 and boiled off repeatedly throughout the part
 of the reaction that required cooling.

Pressure relief was provided at 4 5 the top of the reactor via a one-inch center relief line. A four-inch emergency vent line 6 7 was also at the top of the reactor. This line was closed until the reaction pressure reached 8 9 400 pounds per square inch at which point a 10 rupture disk burst opening the emergency vent 11 line.

12 On the day of the incident the 13 reaction temperature rose uncontrollably likely because cooling was not applied to the 14 15 reactor. The T2 operator reported a cooling problem about ten minutes prior to the 16 explosion. When the team investigated this we 17 found that the cooling system lacked 18 redundancy. This means that a single failure 19 20 of any one of several parts of the cooling 21 system would cause the entire system to fail. 22 Potential failure points included

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1 the drain and supply valves, supply piping 2 blockage, scale build up in the jacket or even a faulty temperature indicator. Because no 3 4 emergency system was immediately available 5 operators would need to manually attach hoses to access an additional water supply in the 6 7 event of a failure. During the incident as the reaction quickly ran away operators would 8 9 not have had the time to perform these 10 necessary actions. CSB investigators found that the 11 12 emergency vent system was designed for normal 13 operation and not designed to consider a runaway reaction. Additionally the set 14 pressure of the emergency vents was set too 15 Had it been lower it would have 16 hiqh. activated before the second exothermic 17 reaction and could have stopped or slowed the 18 runaway by venting at a lower pressure 19 allowing chemicals to boil and remove heat and 20 chemicals from the reactor. 21 22 Although T2 owners conducted

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laboratory testing of the MCMT process
 chemistry, this testing did not look for the
 potential exothermic reaction in the first
 step of the process. The need to cool the
 first step was not identified until the first
 full-scale production batch.

7 Temperature control issues and 8 near miss exothermic reactions occurred in 9 three of the first ten production batches but 10 T2's owners did not systematically investigate 11 these anomalies to determine their cause.

12 Throughout the production history 13 poor yields and batch flow variations occasionally occurred due to temperatures 14 rising too high in the first step. However T2 15 was never able to fully control these batch 16 variations and continued to use the same 17 recipe that it had established in the early 18 batches. 19

Although T2's owners had chemical
industry experience, neither had previously
worked with reactive chemistry and hence did

not have any experience developing and
 designing a reactive exothermic process like
 the MCMT chemistry.

Additionally although T2 was owned 4 5 by a chemist and chemical engineer, neither of them learned about process safety and the 6 7 specific hazards of chemicals of reactive chemistry in their education program. 8 9 Chemistry programs are not 10 intended to address process design and plant production. Chemical engineering programs, 11 while they do cover process and plant design, 12

13 often do not address process safety.

The Mary K. O'Connor Process
Safety Center operated by the Texas A&M
University developed programs to promote
safety in the chemical industry.

18 In 2006 the center conducted a 19 survey of undergraduate chemical engineering 20 programs and found that process safety was a 21 part of the required curriculum in only 11 22 percent of the programs. Within these

specific inclusion of reactive hazard
 recognition and management was even more
 limited.

Recognizing the need for increased 4 5 education and process safety in reactive hazards, the American Institute of Chemical 6 7 Engineers, AIChE, designed and implemented a process safety certificate program which began 8 9 issuing certificates in 2008. The process 10 safety certificate program is free to members. It consists of four modules two of which 11 specifically address reactive hazards. AIChE 12 13 has more than 150 chapters and student membership is free. 14 15 Now to present the root and contributing causes as determined by the team. 16 The root cause of the incident was 17 that T2 did not recognize the runaway reaction 18 hazard associated with the MCMT it was 19 20 producing. As a result the contributing 21 causes were that the reactor cooling system 22 was susceptible to single point failure and it

1 failed on the day of the incident. The 2 reactor relief system was incapable of releasing the pressure from a runaway reaction 3 which resulted in the reactor bursting and the 4 5 contents exploding. Since 1998 the CSB has 6 7 investigated a number of reactive chemical accidents and in 2002 issued a comprehensive 8 9 study of reactive chemical incidents. These 10 investigations and the study included regulatory and industry recommendations aimed 11 at reducing chemical accidents. 12 The 13 recommendations that we propose here today seek to further this work by addressing 14 chemical reactive hazard recognition as part 15 of chemical engineering undergraduate 16 education. 17 Accreditation is a voluntary peer 18 review process to assure quality and specific 19 20 educational programs. The Accreditation Board 21 for Engineering and Technology, or ABET,

22 reviews degree programs in engineering,

1 including chemical engineering. Specialized curriculum criteria are used to review 2 undergraduate chemical engineering programs. 3 The team recommends that the 4 5 American Institute of Chemical Engineers work with ABET to add reactive hazard awareness to 6 7 the required chemical engineering curricula and accredited programs. 8 9 The Process Safety Certificate 10 Program which includes reactive hazard management is offered free to AIChE student 11 12 chapter members. The team recommends that 13 AIChE inform its student members about the process of specific programs and encourages 14 student members to participate. 15 Mr. Bresland, at this point in 16 time the team will entertain any questions 17 that the board has regarding our report. 18 I thank you, 19 CHAIRMAN BRESLAND: Mr. Hall and Mr. Wanko, number one for your 20

22 presentation this evening.

21

diligent work on this and for your excellent

1	We will start on my left with Mr.
2	Wark.
3	MR. WARK: Thank you, Mr.
4	Chairman.
5	I do have a couple questions but I
б	too would like to commend the investigators on
7	an outstanding job. From all accounts this
8	was an outstanding response to a very serious
9	incident and I'd like to also commend the fire
10	community, the other first responders for a
11	job well done as well.
12	I understand that you were delayed
13	in getting onto the site because of the
14	hazardous nature of the materials, the MCMT.
15	How long was it before you could
16	get on the site, if you haven't already
17	mentioned that, and how was the entry made?
18	MR. HALL: After the team arrived
19	we quickly evaluated the hazards associated
20	with the MCMT that T2 was producing. As I
21	mentioned it's a regulated material by the EPA
22	and it's on the list of extremely hazardous

1 substances.

2	As we investigated this we found
3	that there were no commonly available
4	immediate measurement techniques that could be
5	used in the field to detect the MCMT. And we
6	had to go through a lengthy process of taking
7	samples and having them sent to a specialized
8	lab of which there's only one in the country
9	that does this analysis to determine the MCMT
10	levels. And this kept us out of the very core
11	of the site for five weeks.
12	I'd like to also point out that in
13	the emergency response the fire community in
14	entering the site wore protective gear that
15	prevented any injury from this particular
16	highly toxic material.
17	MR. WARK: I have just one more
18	question and that is were there any off site
19	exposures to the chemical?
20	MR. HALL: There were no recorded
21	instances of off site exposure to this
22	particular chemical. The chemical does break

down in sunlight so as it was released it 1 2 broke down in a relatively short period of time so the real danger from exposure to this 3 chemical was within the site, within the 4 5 structures of the site where they were storing the MCMT where sunlight didn't reach. 6 7 MR. WARK: That's all I have, Mr. Chairman. 8 9 CHAIRMAN BRESLAND: Mr. Wright. 10 MR. WRIGHT: Thank you, Mr. Chairman. 11 I too would like to commend all 12 13 those involved with this incident because I think they did an outstanding job. I have 14 several questions for Investigator Hall. 15 One having to do with our scaled down experiment, 16 if you will, to replicate the explosion. You 17 showed a picture of the vessel before and 18 after it under went the secondary thermal 19 reaction or exothermic reaction. 20 My question is did we scale down 21 22 the vents on that particular vessel so it

replicated the same size vents as here? 1 2 That particular test MR. HALL: that was performed was performed in what's 3 known as a closed cell, which is not a vented 4 5 cell. In order to achieve the actual measurements of the heat production and the 6 7 pressure rise, it's a standard practice to do that in a closed cell. 8 MR. WRIGHT: So we didn't have a 9 relief valve? 10 11 MR. HALL: No, there was no relief 12 valve. 13 MR. WRIGHT: And my assumption is 14 because they were unaware of the hazards that they probably didn't conduct any hazard 15 analysis. 16 My question is did you find any 17 evidence of anybody performing any hazard 18 analysis with respect to explosive impacts 19 20 that may occur during the chemical runaway reaction. 21 22 MR. HALL: We did not find any

evidence of a hazard analysis being performed. 1 We did find, however, and it's detailed in the 2 report, that one of the consultants to T2 had 3 4 recommended that they perform a hazard 5 analysis and yet this recommendation was never 6 followed up on. 7 MR. WRIGHT: My final question deals with the -- I believe I read in the 8 9 report that they scaled up production more 10 than once. Initially they started with a small vessel, table top if you will, analysis 11 and then went to full production. 12 13 And then I believe they raised the level of the production again just before this 14 incident? 15 That is correct. 16 MR. HALL: They began their testing with a one-liter reactor, 17 basically a table top device, and did their 18 developmental work in that one-liter reactor 19 20 and went straight to the large reactor that was in the incident. 21 22 For the first 42 batches they used

a recipe that was three quarters the size of 1 2 the recipe that was used the day of the incident. So after batch 42 they began 3 producing a larger quantity, increasing the 4 5 quantity of chemicals in the reactor by one The importance of that is when you do 6 third. 7 that it increases the energy input and makes the situation more dangerous. 8 9 Thank you very much. MR. WRIGHT: 10 CHAIRMAN BRESLAND: Board member Visscher. 11 12 MR. VISSCHER: Thank you, Mr. 13 Chairman. I want to join Mr. Hall and Mr. Wark in mentioning commending the emergency 14 response community and the incident command. 15 They talked about the fact that the response 16 was done in a safe manner, which is the most 17 important thing. And we mentioned, just so 18 you know, the board recently issued a video on 19 20 emergency response and one of the incidents 21 that's held out as an example of good 22 emergency response is the response to this

1 incident.

2	But I also wanted to mention I was
3	a board member who went on at the time of the
4	deployment and wasn't just, although as I
5	said, the most important thing is the safety
6	of the emergency responders. But I was so
7	impressed by the incident command structure
8	and how smoothly all those different agencies
9	were able to work together. And that's to the
10	leadership of the emergency response community
11	and the fire department and the way in which
12	that was handled.
13	So this isn't always the case on
14	our deployments that everything goes so
15	smoothly when you have so many different
16	agencies and local and state and federal
17	agencies involved.
18	We talk about these issues around
19	the country I often use this as one where it
20	really did go very smoothly and again a
21	tribute to the leadership.
22	Question: I gather that the

company and the owners were aware of the first
 exothermic reaction but not the second one in
 that first step. They wanted the first one to
 happen.

5 They needed the first exothermic 6 but they weren't aware of the second one was 7 a possibility?

MR. HALL: That's correct. 8 T mean 9 we did find that actually in their 10 developmental testing when they were testing it in a one-liter reactor they detected no 11 exotherms. And wasn't until the very first 12 13 production batch that they discovered the exotherm, the initial exotherm, in the first 14 They did have occasions where they 15 reaction. 16 allowed the temperature to get too high but in every case they were able to stop this before 17 it proceeded into the uncontrolled part of the 18 second exotherm. 19 Is that also what 20 MR. VISSCHER:

happened, they had previous incidents where it

22 had sort of gotten away?

21

1 MR. HALL: Gotten away but they 2 reestablished cooling and were able to control 3 the incident. 4 MR. VISSCHER: How many of those 5 incidents, do you know? Was that fairly 6 frequent? 7 MR. HALL: WELL, the documentation of that is somewhat limited. All the 8 9 documentation was destroyed. All the on site 10 documentation destroyed and we were able to obtain some back ups from individuals that had 11 documentation off site. But in the initial 12 13 batches, the first ten batches, three of those first ten had exothermic excursions that 14 didn't fully run away but were a cause of 15 16 great concern. 17 MR. VISSCHER: In each case they were able to bring those down by using 18 cooling? 19 MR. HALL: By using cooling to 20 slow it. 21 22 MR. VISSCHER: If you were

starting out a process and you wanted to know 1 what those hazards were, one way to determine 2 what those hazards were would be laboratory 3 testing which they did some of I guess, but 4 5 not as thoroughly perhaps. What would be other -- how would 6 7 you know what the hazards are? THE WITNESS: Well, they did their 8 9 laboratory testing in standard lab equipment. 10 And there's a very specific test process that's used to evaluate exothermic hazards and 11 they refer to it as a thermal hazard 12 13 evaluation test where they used a very specialized device that minimizes heat 14 transfer across the test cell and it more 15 closely behaves like a full-scale reactor. 16 And that was a test device that we had done 17 our testing in that Investigator Wanko had 18 shown that with just 50 grams of the material 19 20 we were able to duplicate a burst vessel like occurred the day of the incident. 21 22 So they had done MR. VISSCHER:

laboratory testing but not necessarily the 1 2 right kind or the laboratory testing that would show -- would indicate those kinds of 3 reactions? 4 5 MR. HALL: That's correct. Mr. VISSCHER: It would seem to me 6 7 that another source of information about hazards would be a literature search. 8 9 Is there literature available in terms of the reactive hazards or the 10 possibilities of an exothermic or runaway 11 reaction of this in the literature? Is there 12 13 a literature search? MR. HALL: In looking at the 14 literature that's available, and they did an 15 extensive literature search in their 16 development. 17 18 This material was developed in the late 1950s by a company called Ethol who got 19 a number of patents for the material. 20 There 21 are some 25 or so patents that are listed in 22 the appendix of the report. These patents

detail how the chemicals were combined to make this product. But the patents don't address hazards or safety aspects of the production. They're purely written to achieve getting the patent from the patent office to give them product protection.

7 Apart from the patent, MCMT from a production standpoint is largely absent from 8 9 the literary universe. There is not much 10 information out there on this particular The information that does exist has 11 material. to do with the hazards and toxicity. 12 13 MR. VISSCHER: Presumably at least now there will be some literature with our 14 report that will highlight some of the hazards 15 to put other people on notice if anyone is 16 thinking of producing it. 17 How did they decide on the rupture 18 disk being set at 400 psig? Was there a 19

20 formula that was used? I think the report

21 indicated that if it had been set

22 significantly lower that would have relieved

the pressure at a much earlier time. 1 MR. HALL: From our understanding, 2 and there was very limited documentation on 3 this. And this particular evidence comes from 4 5 interviews and it's really secondhand information because those that had done this 6 7 work were unfortunately killed in the accident. 8 9 The set point, the size and the 10 set point were based on the normal production of hydrogen from the reaction and also looking 11 at the MAWP of the vessel as well as all three 12 13 steps and what pressure they wanted to operate all three steps of the chemical process at, 14 not really focusing on the first step and 15 looking at the exothermic requirements of the 16 exothermic chemistry. 17 MR. VISSCHER: One last question, 18 Mr. Chairman. 19 The recommendations really focus 20 on improving education in chemical education 21 22 for a better awareness of looking for

1	hazardous reactions or hazards of reactions.
2	Sort of a two-part question.
3	First, why did you focus on education in this
4	report and second of all I think in your
5	introduction the chairman mentioned that your
б	own background is in nuclear engineering. If
7	you would compare sort of the safety awareness
8	in nuclear engineering to the safety awareness
9	in chemical engineering.
10	I take it that what we're doing
11	with this recommendation is urging that the
12	chemical education process, the chemical
13	process, chemical engineering education move
14	in the direction of what's already been done
15	in nuclear engineering.
16	Is that a fair assessment?
17	MR. VISSCHER: That's a fair
18	assessment, Mr. Visscher.
19	When looking at this particular
20	case we saw similarities with a number of
21	other cases where you had a small
22	entrepreneurial company that was going in and

developing a process using a highly hazardous 1 2 reactive chemistry. In these small entrepreneurial companies each one had a 3 4 similarity that there was really a lack of 5 recognition that this hazard existed. When we looked at these small companies, typically 6 7 they were unregulated. They were outside the regulatory environment. The owners were 8 9 unaware of what happened. 10 And we've made regulatory 11 recommendations. We made industry recommendations. We've made recommendations 12 13 that reach out to the big chemical companies and we're really looking for a way as to how 14 can we reach these small companies. And one 15 of the answers is to reach through the 16 education system. It's something we have not 17 done in the multitude of reactive incidents 18 that we've investigated. 19 20 And as you point out, within the 21 nuclear industry there's a big emphasis on 22 education and the learnings that go into the

initial process design and the like in making 1 sure the people fully understand the hazards 2 that they're dealing with and there are checks 3 and balances in the learning aspect of it. 4 5 So in that respect that we had not previously gone this route as a presentation 6 7 measure, we thought that it would be appropriate in this case. 8 9 MR. VISSCHER: Thank you, Mr. Chairman. 10 CHAIRMAN BRESLAND: Other board 11 members for your questions. 12 13 I have a few questions myself. And let me start with I guess it's a follow up 14 to Mr. Visscher's question, maybe in a more 15 16 general sense. We've had this explosion. We've 17 had this tragedy with the four killings. 18 If you were to go back to the 19 20 beginning here, what should have been done that would have prevented this accident from 21 22 happening in terms of starting at the

beginning, developing the process, developing the information, designing the process and operating it?

What could have been done that 4 5 would have prevented this from happening? 6 MR. HALL: Well initially in the 7 developmental stage would be to perform the thermal hazard evaluation testing, to actually 8 9 test for exothermic reactions and get the 10 kinetic, reaction kinetic data to use in the That was kind of the initial 11 process design. point that was missed in the process design. 12 13 Secondly when we look at the --14 those first ten batches that they produced and we saw that they had ten occasions or three 15

occasions in that first ten batches where the 16 temperature went beyond what they had desired 17 and indicating some exothermic activity that 18 may not be desirable, they just kept going 19 20 rather than stopping and trying to evaluate With each one of those they were able 21 that. 22 to catch it. The reactor didn't blow up.

1 They were able to move on. And they

2 normalized that deviation.

And that's much the same way in 3 the Challenger incident, how each time the O 4 5 rings failed, NASA normalized the deviation. The shuttle didn't blow up. We launched again 6 7 and eventually one caught them. And the same thing caught them here. 8 9 So there was multiple opportunities, multiple near misses that 10 occurred throughout the production that were 11 clues to what they should have done. 12 13 CHAIRMAN BRESLAND: So one of the things they should have done was the more 14 sophisticated laboratory testing which I 15 assume they didn't have the capability of 16 17 doing. 18 How many laboratories are there in the United States that have the expertise to 19 20 do this sort of testing? For example how much 21 money did we spend on the testing? 22 MR. HALL: There's about six

1 laboratories in the United States that perform this testing commercially. Also some of the 2 big chemical companies perform the testing 3 internally. But those that perform it as a 4 service there's about six laboratories in the 5 country performing it commercially. 6 7 The CSB spent about \$125,000 to perform this particular test and got some of 8 9 the top people in the country known for this 10 particular type of testing to guide the test and evaluate the results. 11 CHAIRMAN BRESLAND: Who else makes 12 13 this chemical in this country? MR. HALL: There is one other 14 domestic producer and that is a company called 15 It used to be called Ethyl 16 Afton. Corporation. And they have been producing 17 this particular material since the late 1950s. 18 19 CHAIRMAN BRESLAND: Have they had 20 any incidences that we know of? MR. HALL: We found no record of 21 22 incidents that they had had that were publicly

1 known.

2	CHAIRMAN BRESLAND: The vessel
3	exploded at you say about 5000 psi and it was
4	a three-inch vessel so it was a very
5	significant explosion.
б	Assuming that it that the
7	system was designed properly, how big would
8	the rupture disk have to have been? Keeping
9	in mind that the rupture disk is something
10	that as you know is something that blows
11	before the vessel blows so it avoids the
12	catastrophic explosion of the vessel.
13	Do we know, do we have any
14	estimates from any of our experts how big that
15	rupture disk would needed to have been to
16	avoid this catastrophic destruction of the
17	vessel?
18	MR. HALL: From our testing we
19	looked at what relief size. And there's two
20	components to the relief size, the set
21	pressure as well as the physical size. And at
22	400 psi, which was the set pressure chosen by

T2, there was not a size big enough that would 1 2 relieve the pressure and protect the vessel. However, if you drop that set 3 point down to 75 psi, and what this allows to 4 5 happen is when it relieves at the lower pressure the chemicals inside the reactor 6 7 immediately boil. And once they start boiling that removes heat as well as the chemicals 8 9 then vent and it removes the chemicals from So a relief device set at 75 10 the reaction. 11 psi of approximately the size they had would 12 have protected this vessel. But at 400 psi 13 there was no way to relieve the pressure. There was not a device big enough. 14 CHAIRMAN BRESLAND: 15 In your 16 presentation, just changing topics here, you discuss the cooling system which is the -- not 17 the emergency cooling system but the routine 18 cooling temperature control system, but you 19 20 said there was not a back up to that. What would have -- in your 21 22 experience what would a typical back up system 1 look like?

2 Several of the MR. HALL: employees at T2 indicated that they had a 3 secondary source of water. It wasn't hooked 4 5 up. And what we mean by hooked up is we would need a second line into the cooling jacket 6 7 with its own dedicated automatic valve controlled by the control room. And because 8 9 this secondary source that they were talking 10 about just came from a tank, it would also 11 need a pump or a power supply. So you would 12 need a whole duplicate supply system that 13 could feed water to the reactor jacket. CHAIRMAN BRESLAND: I don't have 14 15 any other questions. 16 Do any of the board members? Board member Wark. 17 18 MR. WARK: Thank you, Mr. Just to follow up with Mr. Hall. 19 Chairman. 20 My presumption is that if people were aware and that if the recommendation is 21 22 heeded in the future, people will do a proper

hazard analysis and will be aware of what 1 potential catastrophic events may occur and 2 hopefully will then site their facilities in 3 4 areas to prevent damage to their neighbors or 5 neighboring businesses. Am I correct in that presumption, 6 7 that that's a hope that we have for the future? 8 9 MR. HALL: Yes, sir. 10 MR. WARK: Thank you. MR. HALL: You know, we found with 11 12 the T2 company that there were some hazards 13 that they were aware of. And on those hazards they spent money and they took protecting 14 action. But in this case it was a hazard they 15 were unaware of and hence appropriately did 16 not deal with it. 17 18 CHAIRMAN BRESLAND: Any other questions? 19 That concludes the questions from 20 the board members. We did have a sign up 21 22 sheet as people walked in and I have a copy of

I have the sign up sheet here but 1 it. unfortunately it's blank. So I don't have 2 any, but that doesn't prevent you from asking 3 questions. So if there are members of the 4 5 audience who would like to either make a comment or ask a question, please do so now. 6 7 With the gentleman here. If you would state your name, please, and spell your 8 9 name for the person who is taking the dictation. 10 11 MR. McDONALD: George McDonald. 12 Had that jacket on the vessel been 13 descaled or do you know if they regularly descaled that vessel, the water jacket? 14 15 CHAIRMAN BRESLAND: Let me repeat 16 the question. The question as I understand it was had the vessel been descaled, meaning? 17 MR. OTHER: You have a water 18 jacket around the vessel where you're putting 19 20 the City of Jacksonville water. 21 MR. BRESLAND: You're asking if 22 there was scale inside?

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1	MR. McDONALD: That's correct. It	
2	boils off so you get a limestone deposit	
3	inside the jacket. You could normally remove	
4	that with an acidic wash.	
5	Do you know if that was done?	
6	CHAIRMAN BRESLAND: Let me turn	
7	that question over to Mr. Hall.	
8	MR. HALL: We have no evidence as	
9	to whether or not that was ever done. We	
10	looked very carefully at the vessel and the	
11	jacket after the explosion but the explosion	
12	was of such magnitude that we felt that any	
13	scale that might have been there could have	
14	been taken off through the very energetic	
15	nature of the explosion. And it was listed as	
16	one of potential single point failures for the	
17	cooling system and it's one of the possible	
18	failures that the team considered causing a	
19	failure of the cooling system.	
20	MR. McDONALD: We've got a very	
21	hard water here, 175 without descaling it	
22	would have significantly reduced the ability	

Page 60 1 to cool. 2 That's all I have. 3 CHAIRMAN BRESLAND: We have someone else coming to the microphone. No, we 4 5 don't. He's one of the audio/visual people. Again if you could tell us your 6 7 name and spell it, please. MS. PADRICK: My name is Sandra 8 9 Padrick, last name is P-A-D-R-I-C-K. I'm curious about whether or not 10 you're subject to standard -- but you did talk 11 about the presence of solvents also in and of 12 13 course if there were such they would have been required to conduct a hazard analysis. So . 14 15 . . 16 CHAIRMAN BRESLAND: The question is were they covered by the OSHA process 17 safety standard? 18 19 MR. OTHER: Yes. Mr. Hall. 20 CHAIRMAN BRESLAND: 21 MR. HALL: The first step of this 22 process OSHA did extensive testing following

1 the accident to determine if the first step was not a covered step, would not be covered 2 under process safety management because the 3 mixture in the reactor did not meet the 4 5 flammability requirement. None of the chemicals listed were covered chemicals so the 6 7 flammability requirement was the only requirement they could have met and OSHA 8 9 determined in their testing that it did not 10 meet the flammability requirement. We did find that T2 Laboratories 11 never did any testing to determine wether or 12 13 not they were covered and did not do what OSHA did to make a determination of coverage. 14 And in fact there were several of the employees 15 that indicated that there may have been some 16 belief that they were covered and they were 17 planning at some time to implement some of 18 those OSHA process safety management but they 19 had not at the time of the incident. 20 21 Thank you. MS. PADRICK: 22 CHAIRMAN BRESLAND: Thank you.

MR. WHEELER: My name is Wess		
Wheeler. And one of the questions that I		
have, Mr. Bresland, is as a requirement of		
back up systems, the emphasis and		
recommendations from the CSB actually say		
that, you know, we want to deal with		
education, which I agree with wholeheartedly.		
But one of the other things that can come from		
this board is a recommendation to OSHA or to		
process safety management systems to have a		
required documentation of at least a back up		
system in place, operational and checked out		
prior to any of these pilot plants actually		
going into production.		
What kind of processes or what		
kind of recommendations can be made in that		
endeavor?		
CHAIRMAN BRESLAND: Let me turn		
that over to Mr. Hall and I can get back to it		
as well and maybe some of the board members		

that, you 6 7 education, 8 But one of 9 this board 10 process sa 11 required do system in p 12 13 prior to an going into 14 15 16 kind of re endeavor? 17 18 19 that over 20 as well and maybe some of the board members

21 would like to comment on that.

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2

3

4

5

22 MR. HALL: As I mentioned in the

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presentation, the CSB has investigated a number of reactive chemical incidents over the years since our formation and since we began investigating incidents in 1998 and have made a large number of regulatory and industry recommendations with regard to that.

7 The back up systems are covered as part of standard design practice, it's 8 9 something OSHA refers to as generally accepted 10 engineering practice. So they really do come into the regulations in that regard because 11 there are standards and guidelines out there 12 13 that address how to design the system and they're included in those guidelines and OSHA 14 incorporates those within their regulations. 15 CHAIRMAN BRESLAND: Back almost 16 seven years to the day, I think it was on 17 September 17th, 2002, we did make a 18 recommendation to OSHA on the issue of 19 20 reactive chemicals and that's still an open 21 recommendation. It's a complex 22 recommendation.

1 But if a facility like this, 2 regardless of whether it's covered under reactive chemical or under the facts that it 3 has flammable solvents, there would be a 4 5 standard list of activities that they would be required to comply with. And one of them 6 7 would be looking at doing a process hazard analysis which is what we talked about 8 9 earlier. 10 And in that process analysis I would hope that they would look at the process 11 and decide maybe a single source of emergency 12 13 cooling, a single source of cooling water is not appropriate and they would need a back up 14 source of cooling water that would be 15 controlled separately from the controlled 16 It would be what we call a hard wired 17 system. system which would automatically come on 18 immediately. 19 20 I know I struggle with this whole issue of why did this accident happen. 21 Why 22 did the company not do the appropriate testing Page 64

1 to determine what are the hazards associated with this? Because when we did the testing we 2 were able to find quite quickly that there 3 4 certainly was a hazard when you got to a 5 certain temperature in the reaction. MR. WHEELER: 6 I agree that the 7 back up system in place might have helped prevent this and prevent the loss of 8 9 businesses and the loss of life that we 10 actually had. And that was one of the things that I would have liked to have seen come out 11 of this, some way we could document that the 12 13 back up system is in place like you said hard wired and separate from the automation so that 14 we have that in place before these 15 entrepreneurs actually go out and venture into 16 this realm of an unknown uncertainty. 17 18 Thank you very much. 19 CHAIRMAN BRESLAND: It's certainly 20 not unusual to have a back up system, a back 21 up shut down system, a back up emergency 22 system that is hard wired separate from the

distributive control system that operates the 1 process. I'm not talking about this process 2 but just in the chemical and refining industry 3 4 in general. That there's a particular hazard 5 that's a particularly dangerous hazard, that 6 would be a common practice that would be put 7 in place. 8 MR. WHEELER: Thank you. 9 CHAIRMAN BRESLAND: One more 10 gentleman coming up. 11 MR. McMASTER: Jim McMaster. And I have a question. 12 13 You mentioned that a company, 14 Afton, produces the same product. Have they been investigated or 15 looked over to see if they have in place the 16 needed safety measures to continue operation 17 so this doesn't happen again? 18 CHAIRMAN BRESLAND: Mr. Hall. 19 20 MR. HALL: The CSB did not conduct 21 an investigation of Afton. Afton is, though, 22 a large chemical company that's been operating

for many years and we, you know, look to see 1 from our data, from the public domain if there 2 had ever been any accidents with this material 3 and detected no accidents with this material. 4 5 Our particular charter to investigate here really didn't allow us to go investigate 6 7 Afton's activities after the fact. MR. McMASTER: All right. 8 Thank 9 you. 10 CHAIRMAN BRESLAND: In your conversations with Afton, Mr. Hall, were they 11 aware of the second exotherm that we 12 13 discovered in our testing? MR. HALL: We did have several 14 conversations with Afton and they were fully 15 aware of the behavior of this material. 16 17 CHAIRMAN BRESLAND: Thank you. Any more people who'd like to come 18 up and make a comment or ask a question? 19 20 MR. ROGERS: Good evening. My name is AAron Rogers. You've referred to 21 Afton several times here. 22

1 So if they were aware of the 2 second exothermic reaction, had they conducted a process hazard analysis? 3 Mr. Hall. 4 MR. BRESLAND: 5 MR. HALL: Because we didn't investigate Afton I really don't know the 6 7 answer to that question. MR. ROGERS: Thank you. 8 9 CHAIRMAN BRESLAND: Do we have 10 anyone else who would like --11 MR. SCOPPA: My name is Rick Scoppa, S-C-O-P-P-A. 12 13 You mentioned that there's one other company that makes this material and 14 that this company started making it and it's 15 an extremely hazardous material. I don't 16 understand how they could get started to make 17 this without one of the regulatory agencies 18 being heavily involved with making sure that 19 20 they did do the testing necessary to ensure 21 their process system was accurate or proper. 22 CHAIRMAN BRESLAND: Are you

referring to Afton or are you referring to T2? 1 MR. SCOPPA: 2 т2. 3 CHAIRMAN BRESLAND: Okay. 4 MR. SCOPPA: With it being an 5 extremely hazardous material, I don't understand how a company can get started 6 7 without at least the fire department knowing they were making it so they could have had 8 9 some response drills prior to the incident. 10 CHAIRMAN BRESLAND: Well, I see there are some members of the fire department 11 here this evening and I'm not going to put 12 13 them on the spot. But I know that there are requirements not connected to the actual 14 design construction operation of the facility. 15 But there are requirements that companies 16 report to the fire department and to the local 17 emergency planning committee the chemicals 18 that they have on site so that the appropriate 19 20 emergency response procedures and planning can 21 be put in place.

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I think my understanding is, I'm

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going to ask Mr. Hall this question, was that 1 2 there was a partial reporting but not a complete reporting of the chemicals that were 3 on site at that location. And let me ask Mr. 4 5 Hall to expand on that. 6 MR. HALL: In the investigation we 7 found there's a particular regulation which requires reporting of materials under a tier 8 9 two report to the local emergency planning 10 commission, the fire department and the like. They submitted a report under this 11 particular regulation, it's an EPA regulation, 12 13 that reported the various solvents and the raw materials and the sodium that they had on 14 site. What they failed to include in this 15 16 particular report was the actual methylcyclopentadiene and diethylene glycol 17 which is an extremely hazardous substance and 18 it was not included in the report that was 19 20 sent to the LAPC and the fire department. 21 The fire department was aware of the solvents that were on the site and did 22

inspections. They were aware of the sodium 1 2 that was on site and did inspections. And I believe they actually did drills dealing with 3 a sodium mishap on site. But they were never 4 5 informed of the MCMT production that was going on at the site. The fire department was there 6 7 and looked at the hazards they were told about. 8 9 CHAIRMAN BRESLAND: Thank you. 10 Do we have anyone else who would like to make a comment or ask question? Okay. 11 12 We'll move on to the next phase of 13 the meeting which is when we take a vote on the report and the recommendations. 14 And I'd like to ask if one of the 15 board members would like to make a motion? 16 MR. WARK: Yes, Mr. Chairman, I'd 17 like to make a motion. I'd like to move that 18 the board approve the CSB investigative report 19 and recommendations No. 2008-03-I-FL regarding 20 the agency's investigation into the explosion 21 and chemical fire that occurred on December 22

19, 2007 at the T2 Laboratories, Incorporated 1 in Jacksonville, Florida. 2 3 CHAIRMAN BRESLAND: Thank you Board Member Wark. 4 5 Do we have a seconder? MR. WRIGHT: I second, Mr. 6 7 Chairman. CHAIRMAN BRESLAND: I'd like to 8 9 restate the motion. And the motion is to 10 approve the CSB investigative report and Recommendation No. 2008 dash 03 dash 1 dash 11 Florida, regarding the agency's investigation 12 13 into the explosion and chemical fire that occurred on December 19th, 2007 at the T2 14 Laboratories, Incorporated in Jacksonville, 15 Florida. 16 Do we have any discussion on this 17 motion from the board members? 18 19 Board Member Wark? MR. WARK: No, I have no 20 21 discussion. 22 CHAIRMAN BRESLAND: Board Member

1 Wright? 2 MR. WRIGHT: No. 3 CHAIRMAN BRESLAND: Board Member Visscher? 4 MR. VISSCHER: 5 No. CHAIRMAN BRESLAND: In that case 6 7 we'll close the discussion on this motion and we'll call for a vote and we'll start at the 8 9 far end with Board Member Wark. You either 10 approve or disapprove of the motion. MR. WARK: Approve. 11 12 CHAIRMAN BRESLAND: Board Member 13 Wright? MR. WRIGHT: Mr. Chairman, I 14 15 approve. 16 CHAIRMAN BRESLAND: Board Member Visscher? 17 18 MR. VISSCHER: Approve. 19 CHAIRMAN BRESLAND: And I approve also board members. 20 21 Based on that the motion has been 22 approved and the investigation report and its

recommendations have been approved by the
 board. Which brings us to some closing
 remarks. Just a couple of points I'd like to
 make.

5 The members this evening have made 6 reference to the emergency responders who 7 responded to this incident. To those of you who saw the photographs of the incidents, the 8 9 dramatic photographs of the fire, some of them are here this evening. Let me just ask, I see 10 a lot of firefighters here just a show of 11 hands of firefighters if you did actually 12 13 respond to the incident. 14 (Applause) CHAIRMAN BRESLAND: 15 Any other emergency responders, EMTs, paramedics? Okay. 16 Well, again thank you very much 17 for your response under very dangerous 18 circumstances. 19 I'd also like to point out that 20 21 this morning we had a separate meeting with 22 the family members of some of the people who

were either killed or injured in the accident 1 and also some of -- there were some people 2 there who actually were at the site on the day 3 4 of the explosion and in one case very 5 seriously injured. We were able to give them basically the same presentation that we gave 6 7 here this evening but in a quieter, more sober setting. And a lot of them could get an 8 9 education what actually happened. 10 For me and for Mr. Hall who was 11 there, it was a very sobering experience to meet those people and realize the terrible 12 13 tragedy that impacted them and the fact that they are going to be dealing with this for the 14 rest of their lives, especially the people who 15 16 were very seriously injured. I would like to thank each of the 17 board members for your participation. 18 My thanks go to Mr. Hall and his team for their 19 20 diligence in carrying out this complex and difficult investigation. All of us here have 21 22 a strong interest in preventing these tragic

1 incidents from occurring.

2	Our goal at the Chemical Safety
3	Board is to ensure that accidents do not occur
4	in the future as a result of runaway chemical
5	reactions, and that future chemical
6	engineering students will be educated about
7	the hazards associated with reactive
8	chemistry.
9	In the next few months the CSB
10	will be working with the American Institute of
11	Chemical Engineers and the Accreditation Board
12	to ensure implementation of the safety
13	recommendations that were approved here this
14	evening.
15	Just because we pass a vote and we
16	make recommendations, it doesn't stop here
17	this evening. We continue for several years
18	with follow up on the recommendations to make
19	sure that to the best of our ability that the
20	recommendations are complied with and we will
21	be in communication with AIChE and with the
22	Accreditation Board to make sure that they do

		Page	77
1	move forward in the way that we would		
2	recommend.		
3	I would like to thank all of		
4	today's participants. I'd like to thank the		
5	audience for your attention. I'd like to		
6	thank the people who took the opportunity to		
7	come up and make comments also.		
8	And with that this evening's		
9	meeting is adjourned.		
10	(Whereupon, the proceedings were		
11	adjourned at 7:30 p.m.)		
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15			
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